

Name: _____

Naming Ionic Compounds

1. Use the Periodic Table to find the charges of the monatomic (simple) ions made from atoms of the following elements. Write the charges as superscripts.

- a. Li
- b. Ba
- c. Ga
- d. Br
- e. S
- f. F
- g. Mg
- h. Al
- i. O
- j. K
- k. Cl
- l. Be

2. Name the following simple ions.

- a. Na^+
- b. N^{3-}
- c. Cr^{6+}
- d. H^-
- e. S^{2-}
- f. Al^{3+}

3. Write the proper charges of the following polyatomic ions as superscripts, and give their names.

- a. SO_4
- b. ClO_2
- c. CO_3
- d. HCO_3
- e. IO_4
- f. SH

4. Write the molecular formulas, with proper subscripts, for the following ionic compounds.

- a. iron(III) chloride
- b. calcium sulfide
- c. sodium carbonate
- d. magnesium nitride
- e. cobalt(II) nitrate
- f. sodium phosphate

NAMING IONIC COMPOUNDS WORKSHEET

- g. potassium hydrogen sulfate
- h. aluminum hydroxide
- i. calcium hypochlorite
- j. hydrogen chloride
- k. barium phosphate
- l. arsenic(III) oxide

5. Give the proper names for the following ionic compounds.

- a. BaO
- b. Al₂O₃
- c. Na₂S
- d. BN
- e. FeCl₃
- f. BaSO₄
- g. NaClO
- h. HgCl₂
- i. K₃PO₄
- j. Na₂CO₃
- k. SnCl₂
- l. V₂O₅
- m. Ca(HSO₄)₂
- n. CrO₃